

**Class X Session 2024-25**  
**Subject - Science**  
**Sample Question Paper - 8**

**Time Allowed: 3 hours**

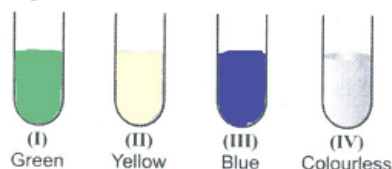
**Maximum Marks: 80**

**General Instructions:**

1. This question paper consists of 39 questions in 5 sections.
2. All questions are compulsory. However, an internal choice is provided in some questions. A student is expected to attempt only one of these questions.
3. Section A consists of 20 objective-type questions carrying 1 mark each.
4. Section B consists of 6 Very Short questions carrying 02 marks each. Answers to these questions should be in the range of 30 to 50 words.
5. Section C consists of 7 Short Answer type questions carrying 03 marks each. Answers to these questions should be in the range of 50 to 80 words.
6. Section D consists of 3 Long Answer type questions carrying 05 marks each. Answers to these questions should be in the range of 80 to 120 words.
7. Section E consists of 3 source-based/case-based units of assessment of 04 marks each with sub-parts.

**Section A**

1. Four test tubes containing solutions (I), (II), (III) and (IV) are shown below along with their colours. Zinc sulphate is contained in [1]



- a) II b) IV
- c) I d) III
2. Choose a displacement reaction: [1]
- a) Burning of metals b) Addition of more active metal to a solution of a less active metal compound.
- c) Extraction of metals d) Electrolysis
3. The compounds used to prepare  $\text{NaHCO}_3$  are: [1]
- a)  $\text{Na}_2\text{CO}_3$ ,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{O}_2$  b)  $\text{NaCl}$ ,  $\text{NH}_3$ ,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$
- c)  $\text{NaCl}$ ,  $\text{NaOH}$ ,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$  d)  $\text{NaCl}$ ,  $\text{Ca(OH)}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{CO}_2$
4. The hardness of water is caused by: [1]

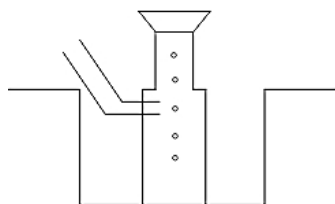
a) All of these

b)  $\text{Mg}(\text{HCO}_3)_2$

c)  $\text{CaCl}_2$

d)  $\text{CaSO}_4$

5. A metal is heated with dil  $\text{H}_2\text{SO}_4$ . The gas evolved is collected by the method shown in the figure. Answer the following questions based on it: [1]



The gas \_\_\_\_\_ than air and it is \_\_\_\_\_ in water.

a) heavier, insoluble

b) lighter, soluble

c) heavier, soluble

d) lighter, insoluble

6. Among the following, the metal with lowest density is: [1]

a) Lead

b) Magnesium

c) Aluminium

d) Lithium

7. The total number of electrons shared in the formation of an ethyne molecule is: [1]

a) 10

b) 4

c) 3

d) 6

8. The process common to aerobic and anaerobic respiration is [1]

a) Glycolysis

b) Electron transport chain

c) Krebs's cycle

d) Oxidation

9. Father of Human genetics is [1]

a) H.G Khurana

b) Sir Archibald Garrod

c) Gregor Mendel

d) Charles Darwin

10. In a flower, the parts that produce male and female gametes (germ cells) are [1]

a) stamen and style

b) filament and stigma

c) anther and ovary

d) stamen and anther

11. Alternative forms of a gene are called [1]

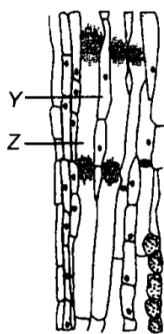
a) Chromosomes

b) Multiples

c) Loci

d) Alleles

12. Refer to the following figure of phloem tissue. Which of the following statements correctly describes the relationship between cells Y and Z? [1]



- a) Y contains ribosomes that synthesise proteins required by Z.
- b) Y provides the carbohydrates that are transported by Z.
- c) Y transports sucrose, while Z transports water and minerals.
- d) Y helps to filter the liquid transported in Z to facilitate translocation.

13. The process of inducing a current in a coil of wire by placing it in a region of changing magnetic field is: [1]

- a) Electrical effect
- b) Magnetic effect of current
- c) Electromagnetic induction
- d) Heating effect of current

14. Match the column I with column II and select the correct option from the codes given. [1]

Column I	Column II
(a) Electric current	(i) volt
(b) emf	(ii) ohm
(c) Resistance	(iii) ohm-metre
(d) Resistivity	(iv) ampere

- a) (a) - (iii), (b) - (i), (c) - (ii), (d) - (iv)
- b) (a) - (iv), (b) - (ii), (c) - (i), (d) - (iii)
- c) (a) - (iv), (b) - (i), (c) - (ii), (d) - (iii)
- d) (a) - (iii), (b) - (iv), (c) - (i), (d) - (ii)

15. A system of inter-dependent food chains represents [1]

- a) Ecosystem
- b) Community
- c) Trophic levels
- d) Food web

16. In the following groups of materials, which group contains only non-biodegradable materials? [1]

- a) Polythene, Detergent, PVC
- b) Wood, Paper, Leather
- c) Plastic, Detergent, Grass
- d) Plastic, Bakelite, Cloth

17. **Assertion (A):** Sodium metal is stored under Kerosene. [1]

**Reason (R):** Metallic sodium melts when exposed to air.

- a) Both A and R are true and R is the correct explanation of A.
- b) Both A and R are true but R is not the correct explanation of A.
- c) A is true but R is false.
- d) A is false but R is true.

18. **Assertion (A):** Asexual reproduction is a primitive type of reproduction. [1]

**Reason (R):** Asexual reproduction involves only mitotic cell division.

- a) Both A and R are true and R is the correct
- b) Both A and R are true but R is not the

explanation of A.

correct explanation of A.

c) A is true but R is false.

d) A is false but R is true.

19. **Assertion (A):** Force experienced by moving charge will be maximum if direction of velocity of charge is perpendicular to applied magnetic field. [1]

**Reason (R):** Force on moving charge is independent of direction of the applied magnetic field.

a) Both A and R are true and R is the correct explanation of A.

b) Both A and R are true but R is not the correct explanation of A.

c) A is true but R is false.

d) A is false but R is true.

20. **Assertion (A):** Ecology is study of relationship between living organisms and their environment. [1]

**Reason (R):** The biotic community and non-living environment of an area function together to form an ecosystem.

a) Both A and R are true and R is the correct explanation of A.

b) Both A and R are true but R is not the correct explanation of A.

c) A is true but R is false.

d) A is false but R is true.

### Section B

21. Classify the following compounds as alkanes, alkenes and alkynes:  $C_2H_4$ ,  $C_3H_4$ ,  $C_4H_8$ ,  $C_5H_{12}$ ,  $C_5H_8$ ,  $C_3H_8$ ,  $C_6H_6$  [2]

22. Why cannot fertilisation take place in flowers if pollination does not occur? [2]

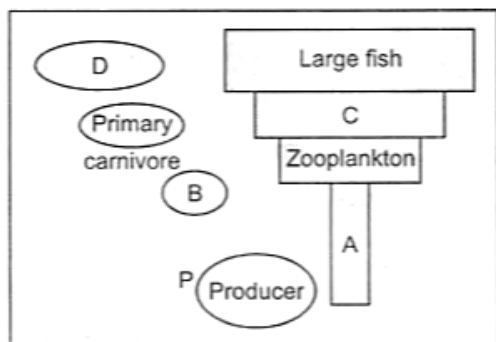
23. What happens to glucose which enters the nephron along with the filtrate? [2]

OR

Explain the role of liver in excretion.

24. If the image formed by a lens for all positions of an object placed in front of it is always erect and diminished, what is the nature of this lens? Draw a ray diagram to justify your answer. [2]

25. Complete the below diagram by filling spaces marked as A, B, C and D: [2]



OR

Consider the following food chains-

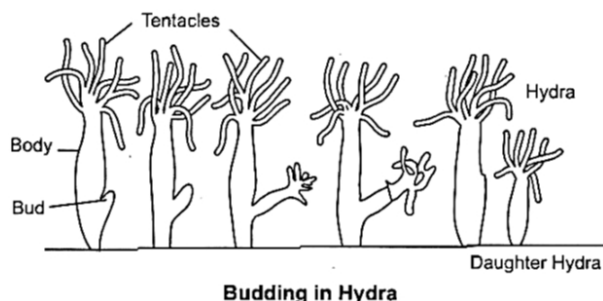
- a. Plants → Mice → Snakes → Hawks
- b. Plants → Mice → Hawks

If energy available at the producer level in both the food chains is 100J. In which case will hawks get more energy and how much & Why?

26. List two main causes of a person developing near sightedness. Show with the help of a ray diagram how this defect can be corrected. [2]

### Section C

27. State three reasons for the following facts: [3]
- Sulphur is a non-metal.
  - Magnesium is a metal.
- One of the reasons must be supported with a chemical equation.
28. [3]
- Define non-metals. Give five examples of non-metals.
  - Name a non-metal which conducts electricity.
  - Name a non-metal having lustre (shining surface).
  - Name a non-metal which is extremely hard.
  - How do non-metals react with oxygen? Explain with an example. Give equation of the reaction involved. What is the nature of the product formed? How will you demonstrate it?
- OR
- Write the electron-dot structures for sodium, oxygen and magnesium.
  - Show the formation of  $\text{Na}_2\text{O}$  and  $\text{MgO}$  by the transfer of electrons.
  - What are the ions present in these compounds?
29. [3]
- Why is nutrition a necessity for an organism? State three reasons.
  - What is likely to happen if green plants disappear from Earth?
30. Explain Mendel's observation when he crossed a homozygous tall (TT) plant with homozygous dwarf (tt) plant followed by self-cross. [3]
31. The refractive indices 1.0003, 1.31 1.5 respectively of Air, Ice and Benzine in which of these does the light travels fastest? [3]
32. [3]
- A wire of resistance 2 has been connected to a source of 50 V as its two ends. What is the current flowing through the wire?
  - An electric refrigerator rated 400 W operates 8 hour/day. What is the cost of the energy to operate it for 30 days at ₹ 3.00 per kWh?
33. A household uses the following electric appliances: [3]
- The refrigerator of rating 400 W for 10 h each day and Two electric fans of rating 80 W each for 6 h daily.
  - Six electric tubes of rating 18 W each for 6 h daily.
- Calculate the electricity bill for the household for the month of June, if the cost of electrical energy is Rs 3 per unit.
- Section D**
34. Esters are sweet-smelling substances and are used in making perfumes. Suggest some activity and the reaction involved for the preparation of an ester with a well-labeled diagram. [5]
- OR
- Define structural isomer and draw the isomeric structures of butane. Compare the structure of benzene and cyclohexane by drawing them.
35. With the help of suitable diagrams explain the various steps of budding in Hydra. [5]



OR

Why do we call pituitary gland as the master gland? Where is it located and what are its functions?

36. i. Draw a labelled ray diagram to show the path of a ray of light incident obliquely on one face of a glass slab. [5]  
 ii. Calculate the refractive index of the material of a glass slab. Given that the speed of light through the glass slab is  $2 \times 10^8$  m/s and in air is  $3 \times 10^8$  m/s.  
 iii. Calculate the focal length of a lens, if its power is - 2.5 D.

OR

- a. Draw a ray diagram for showing the convergence of a parallel beam of light by a concave mirror and mark the positions of pole and the centre of curvature in it.  
 b. An object 4 cm in size is placed at 25 cm in front of a concave mirror of focal length 15 cm. At what distance from the mirror should a screen be placed in order to obtain a sharp image? Find the nature and size of the image.  
 c. List any two uses of a concave mirror.

#### Section E

37. **Read the text carefully and answer the questions:** [4]

Salt of a strong acid and strong base is neutral with a pH value of 7. NaCl common salt is formed by a combination of hydrochloride and sodium hydroxide solution. This is the salt that is used in food. Some salt is called rock salt, bed of rock salt was formed when seas of bygone ages dried up. The common salt thus obtained is an important raw material for various materials of daily use, such as sodium hydroxide, baking soda, washing soda, and bleaching powder.

- (a) If given acids are phosphoric acid, carbonic acid, hydrochloric acid and sulphuric acid, then which acid does not form an acidic salt?  
 (b) What is the formula of baking soda?

OR

Name the substance which on treatment with chlorine to obtain bleaching powder.

38. **Read the text carefully and answer the questions:** [4]

Following questions are based on the two tables given below. Study these tables related to blood sugar levels:

**Table A (Blood glucose chart)**

	Mean Blood Glucose Level (mg/dL)
Doctor's advice needed	380
	350
	315
	280
	250
	215

Good	180
	150
Excellent	115
	80
	50

**Table B (Blood Report of Patient X and Y)**

Time of check	Blood Glucose ranges (mg/dL)	
	Patient X	Patient Y
Before breakfast (Fasting)	<100	70-130
Before lunch, supper and snack	<110	70-130
Two hours after meals	<140	<180
Bedtime	<120	90-150

- Refer to Table B showing the blood report of the levels of glucose of patients X and Y. Infer the disease which can be diagnosed from the given data.
- Identify the hormone whose level in the blood is responsible for the above disease.
- High/low sugar and a low/high-fat diet What would you recommend to the affected patient?

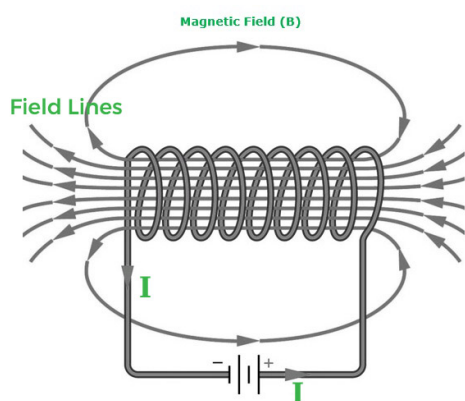
**OR**

Refer to Table A and suggest the value of the mean blood glucose level beyond which doctor's advice is necessary.

39. **Read the text carefully and answer the questions:**

[4]

An insulated copper wire wound on a cylindrical cardboard tube such that its length is greater than its diameter is called a solenoid. When an electric current is passed through the solenoid, it produces a magnetic field around it. The magnetic field produced by a current-carrying solenoid is similar to the magnetic field produced by a bar magnet. The field lines inside the solenoid are in the form of parallel straight lines. The strong magnetic field produced inside a current-carrying solenoid can be used to magnetize a piece of a magnetic material like soft iron when placed inside the solenoid. The strength of the magnetic field produced by a current-carrying solenoid is directly proportional to the number of turns and strength of the current in the solenoid.

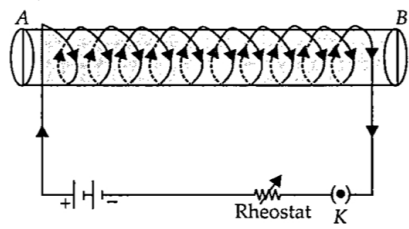


- What would be the strength of the magnetic field inside a long current-carrying straight solenoid?
- Which end is north and which end is south pole when current flows through a solenoid?
- A long solenoid carrying a current produces a magnetic field  $B$  along its axis. If the current is double and

the number of turns per cm is halved, then what will be the new value of the magnetic field?

**OR**

A soft iron bar is enclosed by a coil of insulated copper wire as shown in the figure. When the plug of the key is closed, then where would the face B of the iron bar be marked?






# Solution

## Section A

- (b) IV  
**Explanation:**  $\text{ZnSO}_4$  solution is colourless. It is contained in IV.


- (b) Addition of more active metal to a solution of a less active metal compound.  
**Explanation:** Reactions in which atoms or ions move from one compound to others to form a new compound are known as Displacement reactions.  
A general displacement reaction can be represented using a chemical equation as follows:  
$$\text{A} + \text{BC} \rightarrow \text{AC} + \text{B}$$
  
Displacement reaction takes place only when 'A' is more reactive than B. If 'B' is more reactive than 'A', then 'A' will not displace 'C' from 'BC' and reaction will not be taken place.  
Example: When zinc reacts with hydrochloric acid, it gives hydrogen gas and zinc chloride.
- (a)  $\text{Na}_2\text{CO}_3$ ,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{O}_2$   
**Explanation:** The reaction between sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) and carbon dioxide ( $\text{CO}_2$ ) in the presence of water. Here's the balanced chemical equation for this reaction:  
$$\text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 2\text{NaHCO}_3$$
- (a) All of these  
**Explanation:** The **hardness of water** is caused by magnesium and calcium salts. Calcium and magnesium dissolved in water are the two most common minerals that make water hard. Temporary hardness is a type of water hardness caused by the presence of dissolved bicarbonate minerals (calcium bicarbonate and magnesium bicarbonate).
- (d) lighter, insoluble  
**Explanation:**
  1. Since the gas is collected over water so it is insoluble.
  2. The gas evolved is lighter than air.
- (d) Lithium  
**Explanation:** Lithium has lowest density.
- (a) 10  
**Explanation:** Total number of electrons shared are 10.
- (a) Glycolysis  
**Explanation:** Aerobic respiration has glycolysis continued with the Krebs cycle. In anaerobic respiration, glycolysis ends with lactic acid in humans. Fermentation only has glycolysis with the end product being fermented and transform into acetic acid.
- (c) Gregor Mendel  
**Explanation:** Gregor John Mendel is considered as the father of genetics as he laid down the principles or laws of inheritance for the first time. Though his works were based on plants but the laws governing inheritance patterns are also applicable to humans and hence we call them as "Mendel's Laws of Inheritance".

10. (c) anther and ovary  
**Explanation:**
- The anther is part of the stamen (male sex organ) that produce pollen (male gamete).
  - The ovary is an inferior part of the pistil (female sex organ) which contains ovule. Female gametophyte develops in the ovule.
11. (d) Alleles  
**Explanation:** An alternative form of a gene is known as an allele. Alleles vary in their sequence which may or may not result in a variant phenotype of a particular trait. Alleles represent variations of a gene that is responsible for a particular trait.
12. (a) Y contains ribosomes that synthesise proteins required by Z.  
**Explanation:** In the given figure of phloem tissue. Cell Y is companion cell while Z is sieve tube. Sieve tube lacks cellular organelles, e.g., nucleus, ribosomes, etc. Hence, the adjacent companion cells provide proteins, ATP and other molecules to carry out all the cellular functions in sieve tube cells.
13. (c) Electromagnetic induction  
**Explanation:** Electromagnetic induction
14. (c) (a) - (iv), (b) - (i), (c) - (ii), (d) - (iii)  
**Explanation:** (a) - (iv), (b) - (i), (c) - (ii), (d) - (iii)
15. (d) Food web  
**Explanation:** Food web
16. (a) Polythene, Detergent, PVC  
**Explanation:** Substances that are not broken down into simpler substances by biological processes are said to be non-biodegradable. Polythene, detergents, and PVC (Polyvinylchloride) are non-biodegradable substances. Substances that can be broken down by biological processes are said to be biodegradable. Paper, wood, grass, leather, and cloth are biodegradable.
17. (c) A is true but R is false.  
**Explanation:** A is true but R is false.
18. (a) Both A and R are true and R is the correct explanation of A.  
**Explanation:** Both A and R are true and R is the correct explanation of A.
19. (c) A is true but R is false.  
**Explanation:** From equation  $F = qvB \sin\theta$   
 Force on moving charge will be maximum if the direction of the velocity of charge is perpendicular to direction of magnetic field (when  $\theta = 90^\circ$ )
20. (a) Both A and R are true and R is the correct explanation of A.  
**Explanation:** Both A and R are true and R is the correct explanation of A.

### Section B

21. Alkanes:  $C_5H_{12}$ ,  $C_3H_8$   
 Alkenes:  $C_2H_4$ ,  $C_4H_8$   
 Alkynes:  $C_3H_4$ ,  $C_5H_8$
22. In a flower fertilisation requires both male and female gametes. If pollination does not occur, male gamete is not available hence fertilisation cannot take place.
23. Glucose along with filtrate runs down through the long renal tubule by the action of cilia. Glucose, amino acids, salts, etc., are reabsorbed by the tubular cells and then secreted into the capillary blood cells by diffusion.

OR



Magnesium	Mg	12	2,8, 2	2	Mg:
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ii. Formation of  $\text{Na}_2\text{O}$ :

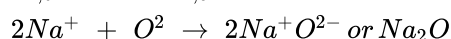
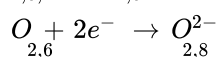
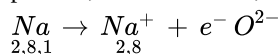
The atomic number of sodium is 11 and it has only one valence electron.

Hence, electronic configuration of  ${}_{11}\text{Na}$  is 2, 8, 1.

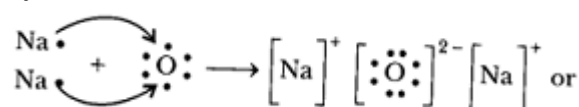
The atomic number of oxygen is 8 and it has 6 electrons in its valence shell.

Hence, electronic configuration of  ${}_{8}\text{O}$  is 2, 6.

Sodium has a tendency to lose the valence electron and oxygen has a tendency to gain the electron lost by sodium. Since, sodium can lose only one electron of the valence shell, and oxygen atom needs two electrons to complete its octet in the valence electron, two atoms of sodium combine with one atom of oxygen. By losing valence electron, sodium is changed into  $\text{Na}^+$  and by gaining two electrons lost by two sodium atoms, oxygen atom is changed into an oxide anion,  $\text{O}^{2-}$ . In this process, both the atoms, sodium and oxygen, obtain the stable electronic configuration of the noble gas neon.



The oppositely charged sodium ion,  $\text{Na}^+\text{O}^{2-}$  and oxide ion,  $\text{O}^{2-}$  are now held together by electrostatic force of attraction or by ionic or electrovalent bond.  $\text{Na}_2\text{O}$  is, therefore, an ionic or electrovalent compound.

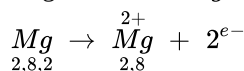


Formation of  $\text{MgO}$ :

The atomic number of magnesium = 12

Its electronic configuration is  $K, L, M$   
 $2, 8, 2$

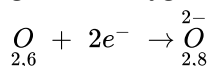
It has two electronic in its outermost shell. So, the magnesium atom donates its 2 valence electrons and forms a stable magnesium ion,  $\text{Mg}^{2+}$  to attain the electronic arrangement of neon atom.



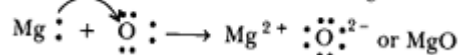
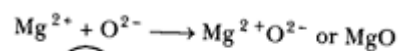
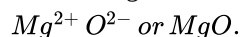
The atomic number of oxygen = 8

Electronic configuration =  $K, L$   
 $2, 6$

It has 6 electrons in its valence shell. Therefore, it requires 2 more electrons to attain the stable electronic arrangement of neon gas. Thus, oxygen accepts 2 electrons donated by magnesium atom and forms a stable oxide ion,  $\text{O}^{2-}$



The oppositely charged magnesium ions,  $\text{Mg}^{2+}$ , and oxide ions, are held together by a strong force of electrostatic attraction to form magnesium oxide compound.



$\text{MgO}$  is ionic compound.

iii. The ions present in  $\text{Na}_2\text{O}$  are sodium ions ( $2\text{Na}^+$ ) and oxide ion  $\text{O}^{2-}$ .

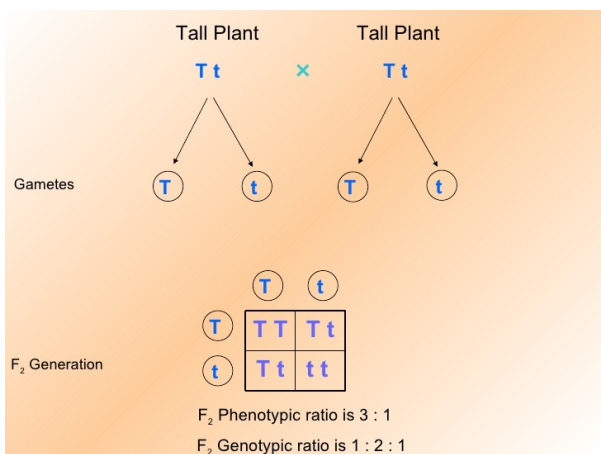
The ions present in  $\text{MgO}$  are magnesium ion ( $\text{Mg}^{2+}$ ) and oxide ion  $\text{O}^{2-}$ .

29. i. Nutrition is necessary because

- it helps in the growth of new cells, survival and maintenance of cells.
- it is needed to develop resistance against diseases.
- it provides energy by the oxidation of food for metabolic processes.

ii. The disappearance of green plants from Earth would mean a total disaster for the ecosystem. It will cause a decrease in concentration of oxygen to such a low level that it would not be sufficient for all the living organism for breathing. Green plants are the source of energy for all organisms. All other organisms directly or indirectly depend on them for food. So, if they disappear from the Earth, all the herbivores will die due to starvation and so will the carnivore. It would result in the extinction of life from the Earth.

30. When Mendel crossed a homozygous tall (TT) plant with homozygous dwarf (tt) plant, all plants in F<sub>1</sub> generation were tall (Tt). Self crossing of F<sub>1</sub> gives F<sub>2</sub>. F<sub>2</sub> generation had 3 tall : 1 recessive plants. Since presence of dwarf allele was masked by tall allele in F<sub>1</sub>, tall allele (T) was dominant over dwarf allele (t).



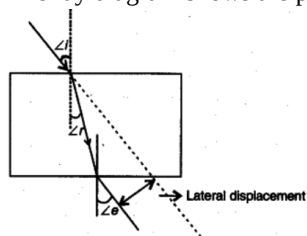
31. Refractive indices of Air, Ice and Benzene are 1.003, 1.31 and 1.5 respectively. Velocity of light in a medium is inversely proportional to refractive index of the medium. Light will travel fastest in air (having least refractive index i.e. 1.0003) and slowest in Benzene (having maximum refractive index i.e. 1.5)
32. i. Zero, current flows due to potential difference and not due to potentials.  
 ii. The total energy consumed by the refrigerator in 30 days would be  
 $400 \text{ W} \times 8.0 \text{ hour/day} \times 30 \text{ days} = 96000 \text{ Wh}$   
 $= 96 \text{ kWh}$   
 Thus the cost of energy to operate the refrigerator for 30 days is  
 $96 \text{ kWh} \times ₹ 3.00 \text{ per kWh} = ₹ 288.00$
33. i. Energy consumed per day by refrigerator  
 $= 0.4 \text{ kW} \times 10 \text{ h} [\because \text{power of refrigerator} = 400 \text{ W} \frac{400}{1000} \text{ kW}] = 0.4 \text{ kWh}$   
 $= 4 \text{ kWh}$   
 Energy consumed per day by fans  
 $= 2 \times 0.08 \text{ kW} \times 6 \text{ h} [\because \text{power of each fan} = \frac{80}{1000} = 0.08 \text{ kW}]$   
 $= 0.96 \text{ kWh}$   
 ii. Energy consumed per day by electric tubes  
 $= 6 \times 0.018 \text{ kWh} \times 6 \text{ h} [\because \text{power of each electric tube} = \frac{18}{1000} = 0.018 \text{ kW}]$   
 $= 0.648 \text{ kWh}$   
 Total energy consumed per day =  $4 + 0.96 + 0.648 = 5.608 \text{ kWh}$   
 Energy consumed in 30 days =  $30 \times 5.608 = 168.24 \text{ kWh}$   
 Cost of 168.24 kWh  
 Cost of 168.24 units @ ₹ 3.00 =  $168.24 \times 3 = ₹ 504.72$

#### Section D

34. i. Take 1 mL ethanol and 1 mL glacial acetic acid along with a few drops of concentrated sulphuric acid in a test tube.  
 ii. Warm the contents in a water bath for at least 5 min.  
 iii. Pour into a beaker containing 20-50 mL of water and smell the resulting mixture.  
 iv. The sweet smell would be observed.



36. i. The ray diagram shows the path of a ray of light incident obliquely on one face of a glass slab:



ii. The glass refractive index is defined as the ratio between the speed of light in the vacuum and the speed of light in the glass.  
Refractive index of glass ( $n_g$ ) = Speed of light in vacuum/speed of light in the glass.

$$n_g = \frac{3 \times 10^8}{2 \times 10^8}$$

$$n_g = 1.5$$

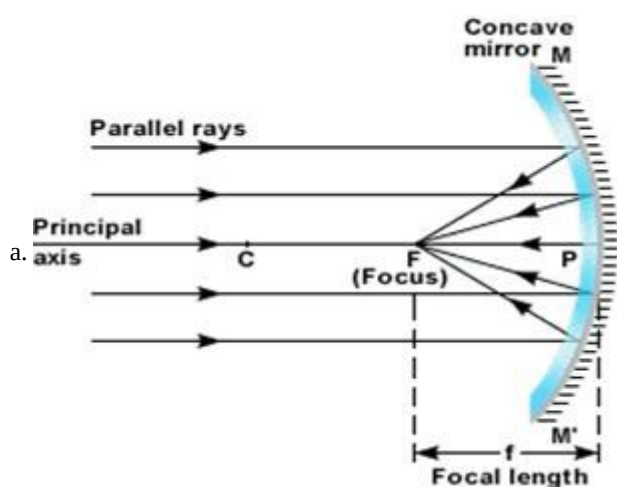
iii.  $P = \frac{1}{f(\text{in meter})}$

$$f = \frac{1}{P} = \frac{1}{2.5} = -\frac{1}{25}$$

$$\Rightarrow f = \frac{-1000}{25} \text{ cm} = -40 \text{ cm}$$

The focal length ( $f$ ) of a concave lens is always negative.

OR



b. Given that  $f = -15 \text{ cm}$ ;  $u = -25 \text{ cm}$ ;  $h_o = 4 \text{ cm}$ ;  $v = ?$

The mirror formula is

$$\frac{1}{f} = \frac{1}{u} + \frac{1}{v}$$

$$\frac{1}{-15} = \frac{1}{-25} + \frac{1}{v}$$

$$\frac{1}{v} = \frac{1}{25} - \frac{1}{15}$$

$$\frac{1}{v} = \frac{3-5}{75} = \frac{-2}{75}$$

$$\Rightarrow v = -37.5 \text{ cm}$$

The screen should be placed 37.5 cm from the pole of mirror and the image is real.

magnification

$$m = \frac{h_i}{h_o} = \frac{-v}{u}$$

$$\frac{h_i}{4} = \frac{-(-37.5)}{-25}$$

$$\Rightarrow h_i = -\frac{37.5}{25} \times 4 = -6 \text{ cm}$$

So the image is enlarged and inverted.

c. Uses of concave mirror:

It is used as a shaving mirror.

It is used as reflector in torch, head light of automobiles etc.

### Section E

37. Read the text carefully and answer the questions:

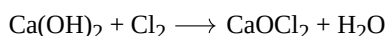
Salt of a strong acid and strong base is neutral with a pH value of 7. NaCl common salt is formed by a combination of hydrochloride and sodium hydroxide solution. This is the salt that is used in food. Some salt is called rock salt, bed of rock salt was formed when seas of bygone ages dried up. The common salt thus obtained is an important raw material for various materials of daily use, such as sodium hydroxide, baking soda, washing soda, and bleaching powder.

(i) Carbonic acid does not form an acidic salt.

(ii) Sodium bicarbonate, commonly known as baking soda or bicarbonate of soda, is a chemical compound with the formula  $\text{NaHCO}_3$ .

OR

$\text{Ca}(\text{OH})_2$  treatment with chlorine to obtain bleaching powder.



**38. Read the text carefully and answer the questions:**

Following questions are based on the two tables given below. Study these tables related to blood sugar levels:

**Table A (Blood glucose chart)**

	Mean Blood Glucose Level (mg/dL)
Doctor's advice needed	380
	350
	315
	280
	250
	215
Good	180
	150
Excellent	115
	80
	50

**Table B (Blood Report of Patient X and Y)**

Time of check	Blood Glucose ranges (mg/dL)	
	Patient X	Patient Y
Before breakfast (Fasting)	<100	70-130
Before lunch, supper and snack	<110	70-130
Two hours after meals	<140	<180
Bedtime	<120	90-15

(i) Diabetes, Diabetes is caused due to less or no secretion of hormone insulin by pancreas.

(ii) Insulin level in the blood is responsible for the given disease.

(iii) Low sugar high fibre diet

OR

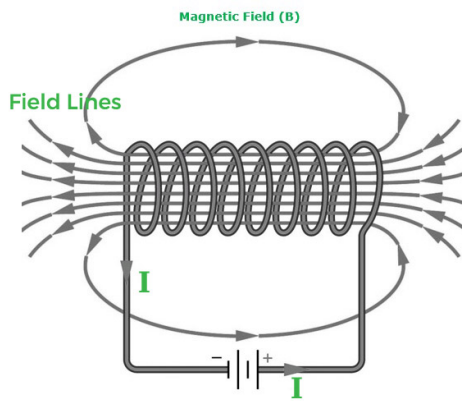
> 180mg/dL.

**39. Read the text carefully and answer the questions:**

An insulated copper wire wound on a cylindrical cardboard tube such that its length is greater than its diameter is called a solenoid. When an electric current is passed through the solenoid, it produces a magnetic field around it. The magnetic field produced by a current-carrying solenoid is similar to the magnetic field produced by a bar magnet. The field lines inside the solenoid are in the form of parallel straight lines. The strong magnetic field produced inside a current-carrying solenoid can be used to magnetize a piece of a magnetic material like soft iron when placed inside the solenoid. The strength of the magnetic field



produced by a current-carrying solenoid is directly proportional to the number of turns and strength of the current in the solenoid.



- (i) Magnetic field inside the infinite solenoid is uniform. Hence it is the same at all points.
- (ii) The end of the current carrying solenoid at which the current flows anti-clockwise behaves as a north pole while that end at which the direction of current clockwise behaves as a south pole and this is according to clock wise.
- (iii) For a long solenoid, magnetic field  $B \propto In$ ; where I is the flowing current and n is number of turns per unit length in the solenoid. Therefore, in the given case magnetic field will remain unchanged.

OR

For a solenoid, if we imagine gripping the solenoid with your right hand so that your curl fingers follow the direction of the current then your thumb will point towards the north end of the electromagnet.